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Question Answer Mark Guidance (ii) n(C. 8. H. 18) burned = 0.32 (mol) 9. 2 (n(CO. 2) from complete combustion) = 2.56 or 2.6 mol OR (ratio n CO. 2 / n C. 8. H. 18) = 7.8(125) OR (n C. 8. H. 18. produce 2.5 mol CO. 2) = 0.31(25) 9. DO NOT ALLOW ECF from an incorrect moles of octane DO NOT ALLOW ECF from incorrect ratio from equation in ALLOW

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(iii) IF answer = 9.34 OR 9.33 % award 3 marks ----- actual  $n(\text{C}_6\text{H}_{13}\text{Br})$   
produced =  $2.31 \text{ g} / 164.9 = 0.0140 \text{ (mol)}$   
theoretical  $n(\text{C}_6\text{H}_{13}\text{Br}) = n(\text{C}_6\text{H}_{14})$   
=  $12.90 \text{ g} / 86.0 = 0.15(0) \text{ (mol)}$  Theoretical  
mass % yield = % yield =  $0.0140 / 0.150$   
 $0.014008$

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Names including alkyl groups: ALLOW  
alkanyl, e.g. ethanyl (i.e. IGNORE 'an')  
DO NOT ALLOW alkol, e.g. ethol (ie 'an'  
is essential) Names of esters: Two words  
are expected, e.g. ethyl ethanoate  
ALLOW one word, e.g. ethylethanoate  
Names with multiple numbers and  
hyphens: Use of 'e'

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