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Chemistry - Chapter 11 - Stoichiometry. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. heididunne. Terms in this set (15) Stoichiometry. the study of the Quantitative, or measurable relationships that exist in chemical formulas and chemical reactions. the Law of Conservation of mass. Stoichiometry is based on _____

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Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical reaction.

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368 Chapter 11 • Stoichiometry Section 11.1.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

Chapter 11: Stoichiometry

In Section 11.3, for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water).

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• Stoichiometry: The quantitative study of the reactants required or the products formed is called stoichiometry. Using stoichiometric calculations, the amounts of one or more reactants required to produce a particular amount of product can be determined and vice-versa. Class 11 Chemistry Notes

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